

Mass and Mole – Chapter 3

If you combust 5.00 g of propane (C_3H_8), how much H_2O (in grams) will you produce?

HINTS:

1. Write and balance the equation for the combustion of propane.
2. Determine the moles of propane in 5.00 g of propane.
3. Determine the number of moles of H_2O you will produce from the moles of propane determined in #2.
4. Determine the grams of H_2O from the moles of H_2O you determined in 3.